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Fundamentals of Analytical Chemistry: 9th ed Chapter 4 Chapter 4 4-1 (a) The millimole is an amount of a chemical species, such as an atom, an ion, a molecule or an electron There are 602 1023 particles 10 3 mole 602 1020 particles mole millimole millimole $\times \times = \times -$ (c) The millimolar mass is the mass in grams of one millimole of a chemical species

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Fundamentals of Analytical Chemistry: 9th ed Chapter 5 (2) After weighing an empty vial, fingerprints are placed on the vial while adding sample to the vial (3) A hygroscopic sample absorbs water from the atmosphere while placing it in a weighing vial 5-5

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An excess of standard AgNO₃ is then added, which precipitates the 4 Fundamentals of Analytical Chemistry: 9th ed Chapter 17 excess tetraphenylborate ion The excess AgNO₃ is then titrated with a standard solution of SCN⁻ The reactions are K⁺ + B(C₆H₅)₄⁻ → KB(C₆H₅)₄ (s) (PDF) Solutions Manual Fundamentals of Analytical

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